26. What water hardness is eliminated by this reaction
$2NaAlSiO_4+CaCl_2=Ca(AlSiO_4)_2+2NaCl$?
A) temporary;
B) carbonate;
C) resistant;
D) average;
E) general.

27. Cation exchange resin to eliminate water hardness corresponds to the formula:
A) $[Cat^+]An^-$;
B) $[An^-]\ Cat^+$;
C) [CatAn];
D) $[CatAn]An^-$;
E) H [Cat].

28. When boiling water, the following reaction occurs:
A) $Ca(HCO_3)_2 \rightarrow CaCO_3+H_2O+CO_2;$
B) $Ca(HCO_3)_2+Ca(OH)_2 \rightarrow CaCO_3+2H_2O;$
C) $Ca(OCI)_2+CO_2+H_2O \rightarrow CaCO_3+2\,HCIO;$
D) $CaCI_2+Na_2CO_3 \rightarrow 2\,NaCI+CaCO_3$;
E) $Ca(HCO_3)_2 \rightarrow CaO+H_2O+CO_2.$

29. The constant water hardness is due to the content in water:
A) calcium and magnesium bicarbonates;
B) bicarbonates and sulfates of calcium and magnesium;
C) chlorides, sulfates, sodium nitrates;
D) chlorides, sulfates, calcium and magnesium nitrates;
E) sodium and potassium bicarbonates.

30. The temporary hardness of the water is due to the content in the water:
A) calcium and magnesium bicarbonates;
B) sodium and potassium bicarbonates;
C) chlorides, sulfates, calcium and magnesium nitrates;
D) chlorides, sulfates, sodium nitrates;
E) bicarbonates and chlorides of calcium and magnesium.

31. Water hardness is determined by:
A) the content of salts of calcium and magnesium;
B) the content of alkali metal sulfates;
C) chloride content;
D) borate content;
E) the content of nitrates and sulfates of metals.

32. The equilibrium constant of the nitrogen oxidation reaction by air depends on:
A) temperature;
B) the concentration of nitrogen;